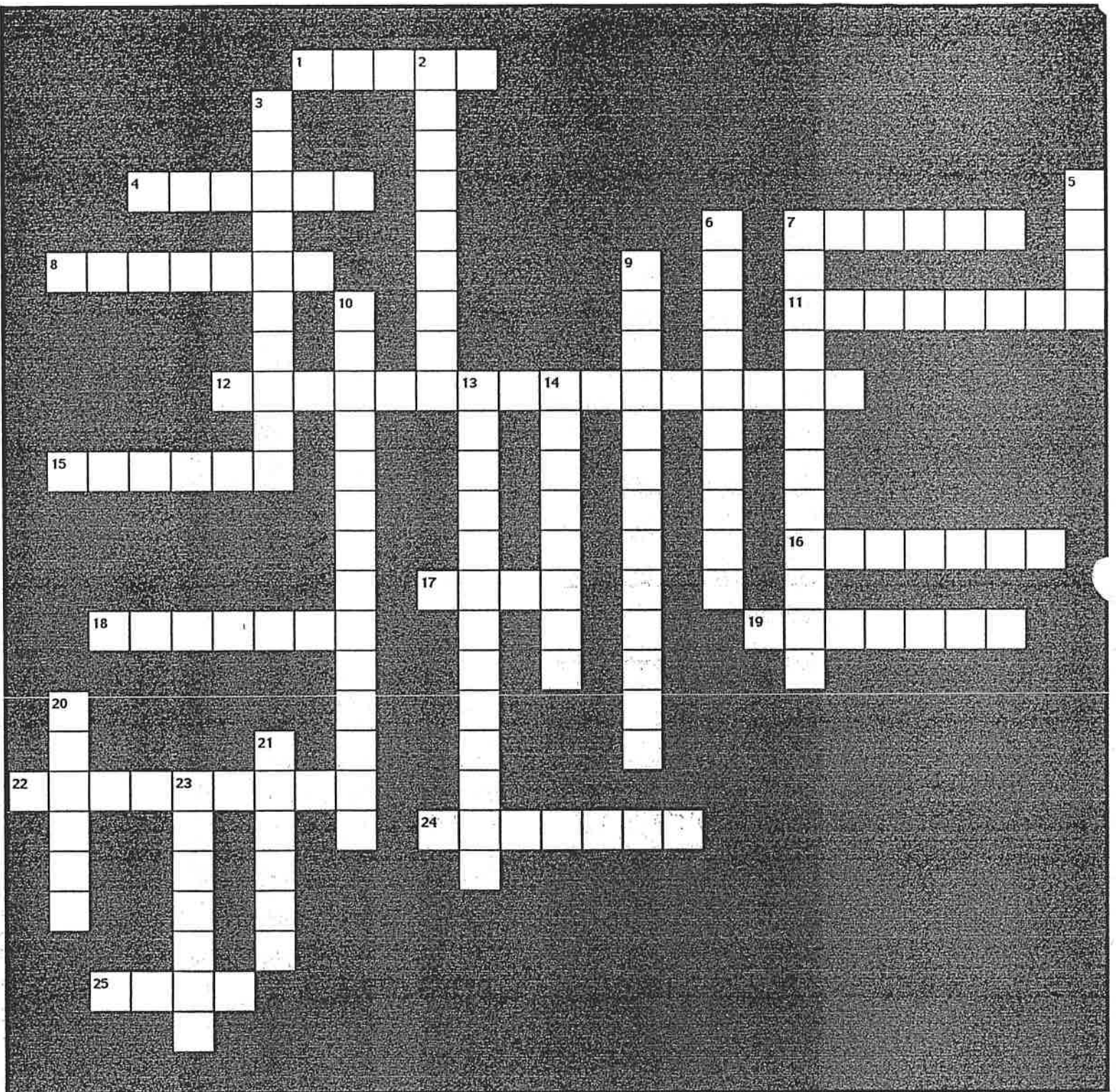


Section 1 Chemistry

Name: _____



Across

1. A positively charged electrode. pg 10
4. The _____ is a positively charged particle present in the nucleus of all atoms. pg 14
7. The number of electrons in a neutral atom _____ the number of protons in the atom. 15
3. This scientist used a glass tube with a small amount of gas and a battery to discover cathode rays. p. 10 fig 4
11. Thompson pictured a sphere of _____ charge with negatively charged electrons spread throughout in his atomic model. p. 12
12. Cathode ray tubes are used in TV screens and _____. p. 10
15. A model is not accepted in the scientific community until it is _____ and the tests support previous observations. p. 12
16. An electrically neutral particle that has the same mass as a proton and is found in an atom's nucleus. p. 15
17. T or F The biggest surprise that came from Thompson's experiments is that there are particles smaller than the atom. p. 11
18. A substance that can't be broken down into simpler substances. p. 9 91
19. A negatively charged electrode. p. 10 91
22. A piece of metal that can conduct electricity. p. 10 1st Sent
24. Most of the mass of an atom is located in its _____. p. 14 92
25. All matter is made up of particles called _____ which means cannot be divided. p. 8 92

Down

2. Dalton proposed that different elements are made of _____ types of atoms. p. 9 92
3. Concluded that atoms are mainly empty space w/ almost all their mass in the nucleus. p. 14 92
5. A proton and a neutron have almost the _____ mass. p. 15 92
6. Thompson concluded that cathode rays are _____ charged particles of matter. p. 11 92
7. Today, scientists will not accept a theory that is not supported by _____ evidence. p. 8 92
9. Theory held by physicists that electrons-like light-have a wave/particle nature. p. 16
10. Rutherford's 1st hypothesis failed when some of the _____ bounced back or veered from the path. p. 13 92
13. Region surrounding the nucleus of an atom, where the electrons are most likely to be found. 17
14. The rest of each atom is occupied by the atom's almost _____ electrons. p. 14 92
20. This scientist pictured an atom as a hard sphere that was the same throughout, like a tiny marble. p. 9 92
21. Thompson's image of an atom was similar to a chocolate chip _____.
23. _____ concluded that an atom is a neutral sphere with negatively charged electrons spread evenly throughout. p. 12 92